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Calculate the molality of each of the following solutions: a. 2.89 g of NaCl dissolved in 0.159 L of water (density of water is 1.00 g/mL) 0.311 molal NaCl b. 1.80 mol KCl in 16.0 mol of H<sub>2</sub>O 6.25 molal KCl c. 13.0 g benzene, C<sub>6</sub>H<sub>6</sub> in 17.0 g CCl<sub>4</sub> 9.80 molal C<sub>6</sub>H<sub>6</sub>

[15.03 - Solution Concentration - Molality, Mass Percent ...](#)

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The following are key terms relevant to the science of osmolality and will help you to develop a better understanding of osmometry: Avogadro's Number: Number of molecules in one mole (gram molecular weight) of a substance. One mole of non-ionic solute (such as sucrose) dissolved in one kilogram of water will yield Avogadro's number (6.02 ...